

2023-24

SCIENCE (033)

CLASS- IX

TOTAL MARKS- 100 { 80 (Theory) + 20 (Internal Assessment)}

THEORY: 80 Marks

Time: 3:00 Hrs.

Unit No.	Unit	Marks
I	Matter- Its Nature and Behaviour	25
II	Organization in the Living World	22
III	Motion, Force and Work	27
IV	Food; Food Production	06
Total		80

Theme: Materials

Unit I: Matter- Nature and Behaviour

Definition of matter; solid, liquid and gas; characteristics - shape, volume, density; change of state- melting (absorption of heat), freezing, evaporation (cooling by evaporation), condensation, sublimation.

Nature of matter: Elements, compounds and mixtures. Heterogeneous and homogenous mixtures, colloids and suspensions. Physical and chemical changes.

Particle nature and their basic units: Atoms and molecules, Law of Chemical Combination, Chemical formula of common compounds, Atomic and molecular masses.

Structure of Atoms: Electrons, protons and neutrons, Valency, Atomic Number and Mass Number, Isotopes and Isobars.

Theme: The World of the Living

Unit II: Organization in the Living World

Cell- Basic Unit of life : Cell as a basic unit of life; prokaryotic and eukaryotic cells, multi-cellular organisms; cell membrane and cell wall, cell organelles and cell inclusions; chloroplast, mitochondria, vacuoles, endoplasmic reticulum, Golgi apparatus; nucleus, chromosomes - basic structure, number.

Tissues, Organs, Organ System, Organism:

Structure and functions of animal and plant tissues (only four types of tissues in animals; Meristematic and Permanent tissues in plants).

Theme: Moving Things, People and Ideas

Unit III: Motion, Force and Work

Motion: Distance and displacement, velocity; uniform and non-uniform motion along a straight line; acceleration, distance-time and velocity-time graphs for uniform motion and uniformly accelerated motion, elementary idea of uniform circular motion.

Force and Newton's laws : Force and Motion, Newton's Laws of Motion, Action and Reaction forces, Inertia of a body, Inertia and mass, Momentum, Force and Acceleration.

Gravitation: Gravitation; Universal Law of Gravitation, Force of Gravitation of the earth (gravity), Acceleration due to Gravity; Mass and Weight; Free fall.

Floatation: Thrust and Pressure. Archimedes' principle; Buoyancy.

Work, Energy and Power: Work done by a Force, Energy, power; Kinetic and Potential energy; Law of conservation of energy.

Sound: Nature of sound and its propagation in various media, speed of sound, range of hearing in humans; ultrasound; reflection of sound; echo.

Theme: Food

Unit IV: Food Production

Plant and animal breeding and selection for quality improvement and management; Use of fertilizers and manures; Protection from pests and diseases; Organic farming.

INTERNAL ASSESSMENT :

20 Marks

- | | |
|--------------------------------------|----------|
| 1- Physics Practical | 03 Marks |
| 2- Chemistry Practical | 03 Marks |
| 3- Biology Practical | 03 Marks |
| 4- Sessional Work | 03 Marks |
| 5- Viva Voce | 03 Marks |
| 6- Continuous Assessment (Unit Test) | 05 Marks |

(There will be total 4 Unit Tests to be conducted throughout the year (two Unit Tests before half yearly examination and two after half yearly examination). At the time of half yearly result preparation best of two Unit Tests (I & II) marks will be taken and converted to the weightage of 05 marks. Likewise best of two Unit Tests (III & IV) marks will be taken and converted to the weightage of 05 marks for the annual result preparation.)

PRACTICALS

Practicals should be conducted alongside the concepts taught in theory classes.

(LIST OF EXPERIMENTS)

- | | |
|---|-----------------|
| 1. Preparation of: | Unit-I |
| a) a true solution of common salt, sugar and alum | |
| b) a suspension of soil, chalk powder and fine sand in water | |
| c) a colloidal solution of starch in water and egg albumin/milk in water and distinguish between these on the basis of | |
| • transparency | |
| • filtration criterion | |
| • stability | |
| 2. Preparation of | Unit-I |
| a) A mixture | |
| b) A compound | |
| using iron filings and sulphur powder and distinguishing between these on the basis of: | |
| (i) appearance, i.e., homogeneity and heterogeneity | |
| (ii) behaviour towards a magnet | |
| (iii) behaviour towards carbon disulphide as a solvent | |
| (iv) effect of heat | |
| 3. Perform the following reactions and classify them as physical or chemical changes: | Unit-I |
| a) Iron with copper sulphate solution in water | |
| b) Burning of magnesium ribbon in air | |
| c) Zinc with dilute sulphuric acid | |
| d) Heating of copper sulphate crystals | |
| e) Sodium sulphate with barium chloride in the form of their solutions in water | |
| 4. Preparation of stained temporary mounts of (a) onion peel, (b) human cheek cells & to record observations and draw their labeled diagrams. | Unit-II |
| 5. Identification of Parenchyma, Collenchyma and Sclerenchyma tissues in plants, striped, smooth and cardiac muscle fibers and nerve cells in animals, from prepared slides. Draw their labeled diagrams. | Unit-II |
| 6. Determination of the melting point of ice and the boiling point of water. | Unit-I |
| 7. Verification of the Laws of reflection of sound. | Unit-III |
| 8. Determination of the density of solid (denser than water) by using a spring balance and a measuring cylinder. | Unit-III |
| 9. Establishing the relation between the loss in weight of a solid when fully immersed in- | Unit-III |
| a) Tap water | |
| b) Strongly salty water with the weight of water displaced by it by taking at least two different solids. | |
| 10. Determination of the speed of a pulse propagated through a stretched string/slinky (helical spring). | Unit-III |
| 11. Verification of the law of conservation of mass in a chemical reaction. | Unit-III |

CLASS- X

TOTAL MARKS- 100 { 80 (Theory) + 20 (Internal Assessment)}

THEORY: 80 Marks

Time: 3:00 Hrs.

Unit No.	Unit	Marks
I	Chemical Substances-Nature and Behaviour	25
II	World of Living	25
III	Natural Phenomena	12
IV	Effects of Current	13
V	Natural Resources	05
Total		80

Theme: Materials

Unit I: Chemical Substances- Nature and Behaviour

Chemical reactions: Chemical equation, Balanced chemical equation, implications of a balanced chemical equation, types of chemical reactions: combination, decomposition, displacement, double displacement, precipitation, endothermic exothermic reactions, oxidation and reduction.

Acids, bases and salts: Their definitions in terms of furnishing of H^+ and OH^- ions, General properties, examples and uses, neutralization, concept of pH scale (Definition relating to logarithm not required), importance of pH in everyday life; preparation and uses of Sodium Hydroxide, Bleaching powder, Baking soda, Washing soda and Plaster of Paris.

Metals and Non-metals: Properties of metals and non-metals; Reactivity series; Formation and properties of ionic compounds; Basic metallurgical processes; Corrosion and its prevention.

Carbon and its compounds: Covalent bonding in carbon compounds. Versatile nature of carbon. Homologous series. Nomenclature of carbon compounds containing functional groups (halogens, alcohol, ketones, aldehydes, alkanes and alkynes), difference between saturated hydro carbons and unsaturated hydrocarbons. Chemical properties of carbon compounds (combustion, oxidation, addition and substitution reaction). Ethanol and Ethanoic acid (only properties and uses), soaps and detergents.

Theme: The World of the Living

Unit II: World of Living

Life processes: 'Living Being'. Basic concept of nutrition, respiration, transport and excretion in plants and animals.

Control and Coordination: Tropic movements in plants; Introduction of plant hormones; Control and co-ordination in animals: Nervous system; Voluntary, involuntary and reflex action; Chemical co-ordination: animal hormones.

Reproduction: Reproduction in animals and plants (asexual and sexual) reproductive health- need and methods of family planning. Safe sex vs HIV/AIDS. Child bearing and women's health.

Heredity: Heredity; Mendel's contribution- Laws for inheritance of traits: Sex determination: brief introduction:

Theme: Natural Phenomena

Unit III: Natural Phenomena

Reflection of light by curved surfaces; Images formed by spherical mirrors, centre of curvature, principal axis, principal focus, focal length, mirror formula (Derivation not required), magnification.

Refraction; Laws of refraction, refractive index.

Refraction of light by spherical lens; Image formed by spherical lenses; Lens formula (Derivation not required); Magnification. Power of a lens.

Functioning of a lens in human eye, defects of vision and their corrections, applications of spherical mirrors and lenses.

Refraction of light through a prism, dispersion of light, scattering of light, applications in daily life.

Theme: How Things Work

Unit IV: Effects of Current

Electric current, potential difference and electric current. Ohm's law; Resistance, Resistivity, Factors on which the resistance of a conductor depends. Series combination of resistors, parallel combination of resistors and its applications in daily life. Heating effect of electric current and its applications in daily life. Electric power, Interrelation between P, V, I and R.

Magnetic effects of current: Magnetic field, field lines, field due to a current carrying conductor, field due to current carrying coil or solenoid; Force on current carrying conductor, Fleming's Left Hand Rule, Direct current. Alternating current: frequency of AC. Advantage of AC over DC. Domestic electric circuits.

Theme: Natural Resources

Unit V: Natural Resources

Our environment: Eco-system, Environmental problems, Ozone depletion, waste production and their solutions. Biodegradable and non-biodegradable substances.

INTERNAL ASSESSMENT :

20 Marks

- | | |
|--------------------------------------|----------|
| 1- Physics Practical | 03 Marks |
| 2- Chemistry Practical | 03 Marks |
| 3- Biology Practical | 03 Marks |
| 4- Session Work | 03 Marks |
| 5- Viva Voce | 03 Marks |
| 6- Continuous Assessment (Unit Test) | 05 Marks |

(There will be total 3 Unit Tests (two Unit Tests before half yearly examination and one after half yearly examination) and a pre-board examination to be conducted throughout the year. At the time of half yearly result preparation best of two Unit Tests (I & II) marks will be taken and converted to the weightage of 05 marks. In annual board examination, marks of the best out of 3 Unit Tests will be taken and converted to the weightage of 05 marks for the result preparation.)

PRACTICALS

Practical should be conducted alongside the concepts taught in theory classes.

LIST OF EXPERIMENTS

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|---|---------------|
| 1. A. Finding the pH of the following samples by using pH paper/universal indicator: | Unit-I |
| (i) Dilute Hydrochloric Acid | |
| (ii) Dilute NaOH solution | |
| (iii) Dilute Ethanoic Acid solution | |
| (iv) Lemon juice | |
| (v) Water | |
| (vi) Dilute Hydrogen Carbonate solution | |
| B. Studying the properties of acids and bases (HCl & NaOH) on the basis of their reaction with: | Unit-I |
| a) Litmus solution (Blue/Red) | |
| b) Zinc metal | |
| c) Solid sodium carbonate | |
| 2. Performing and observing the following reactions and classifying them into: | Unit-I |
| A. Combination reaction | |
| B. Decomposition reaction | |
| C. Displacement reaction | |
| D. Double displacement reaction | |
| (i) Action of water on quicklime | |
| (ii) Action of heat on ferrous sulphate crystals | |
| (iii) Iron nails kept in copper sulphate solution | |

- (iv) Reaction between sodium sulphate and barium chloride solutions
3. Observing the action of Zn, Fe, Cu and Al metals on the following salt solutions: **Unit-I**
- i) $\text{ZnSO}_4(\text{aq})$
 - ii) $\text{FeSO}_4(\text{aq})$
 - iii) $\text{CuSO}_4(\text{aq})$
 - iv) $\text{Al}_2(\text{SO}_4)_3(\text{aq})$
- Arranging Zn, Fe, Cu and Al (metals) in the decreasing order of reactivity based on the above result.
4. Studying the dependence of potential difference (V) across a resistor on the current (I) passing through it and determine its resistance. Also plotting a graph between V and I. **Unit-IV**
5. Determination of the equivalent resistance of two resistors when connected in series and parallel. **Unit-IV**
6. Preparing a temporary mount of a leaf peel to show stomata. **Unit- II**
7. Experimentally show that carbon dioxide is given out during respiration. **Unit-II**
8. Study of the following properties of acetic acid (ethanoic acid): **Unit- I**
- i) Odour
 - ii) solubility in water
 - iii) effect on litmus
 - iv) reaction with Sodium Hydrogen Carbonate
9. Study of the comparative cleaning capacity of a sample of soap in soft and hard water. **Unit- I**
10. Determination of the focal length of: **Unit-III**
- i) Concave mirror
 - ii) Convex lens
- by obtaining the image of a distant object.
11. Tracing the path of a ray of light passing through a rectangular glass slab for different angles of incidence. Measure the angle of incidence, angle of refraction, angle of emergence and interpret the result. **Unit - III**
12. Studying (a) binary fission in *Amoeba*, and (b) budding in yeast and Hydra with the help of prepared slides. **Unit-II**
13. Tracing the path of the rays of light through a glass prism. **Unit-III**
14. Identification of the different parts of an embryo of a dicot seed (Pea, gram or red kidney bean). **Unit-II**
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PRESCRIBED BOOKS:

- Science-Textbook for class IX-NCERT Publication
- Science-Text book for class X- NCERT Publication
- Laboratory Manual-Science-Class IX, NCERT Publication
- Laboratory Manual-Science-Class X, NCERT Publication
- Exemplar Problems Class IX – NCERT Publication
- Exemplar Problems Class X – NCERT Publication
